

# Experimental Determination of the MnO Activity Coefficient in High-Manganese Slags Using the Chemical Equilibrium Method

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**Abstract:** The thermodynamics of manganese oxide in high-MnO-containing slags was investigated using the chemical equilibrium method in the temperature range of 1623 to 1723 K. MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> slags were brought into equilibrium with molten silver (Ag) under controlled CO/CO<sub>2</sub> gas atmosphere. The equilibrium Mn concentration in the silver was measured by ICP-AES (inductively coupled plasma atomic emission spectroscopy) analysis after the experiment. Slag samples were analyzed by EPMA (electron probe microanalyzer) analysis. The obtained activity  $a_{\text{MnO}}$  and activity coefficient  $\gamma_{\text{MnO}}$  were derived as a function of the slag composition and temperature. The activity coefficient of MnO within the investigated slag system increased with an increasing MnO/SiO<sub>2</sub> ratio. The derived temperature dependence of the activity coefficient and partition ratio of Mn between the metal and the slag was strongly influenced by the slag composition. The thermodynamic assessment of the activity and activity coefficient of MnO was carried out by applying the regular solution model (RSM) on the basis of interaction energies of the cations and with FactSage™ 7.3. The theoretical calculations were compared with the experimentally derived values.

**Keywords:** MnO activity; chemical equilibrium; selective reduction; activity coefficient of MnO; manganese distribution; regular solution model

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## 1. Introduction

Manganese is an essential alloying element used during the production of high-grade steel because of its properties, and it is present in most commercially made steels. Mn reduces embrittlement caused by sulfur and improves the mechanical properties of steels, e.g., strength and hardness. Generally, Mn is added as a ferromanganese or silicomanganese alloy in the steelmaking process. Special types of steels require very low concentrations of impurities for their exceptional quality. Therefore, impurity content in the alloy, especially the phosphorus concentration, is absolutely critical, since phosphorus in steel separates at the grain boundaries and causes embrittlement [1–5].

Manganese alloys are made by reducing manganese ore composed of manganese, iron, and silicon oxides as well as phosphorus. During this reduction process, phosphorus oxides are reduced along with manganese oxides, and consequently, phosphorus is transferred into the alloy. Hence, it is necessary to decrease the P content in raw materials, i.e., in the ferromanganese or manganese ore.

Since manganese and phosphorus in steel exhibit similar thermodynamic properties, e.g., both are simultaneously oxidized, it is difficult to remove phosphorus after adding

manganese [4,5]. There have been various studies dealing with the dephosphorization of manganese alloys and manganese ores [6–12].

Fujita et al. [6] proposed a dephosphorization method of the Mn alloy using a flux containing carbonates of basic oxides such as  $\text{BaCO}_3$ . However, this method is considered fundamentally difficult and unfeasible due to many limitations.

Hils et al. [7] proposed a dephosphorization method in the early stages of an integrated process of silicomanganese alloy production. The authors suggested a process of carbon reduction and consequent separation of phosphorus into the metal phase from a mixture of manganese ore and iron ore. It was reported that the phosphorus concentration in the manganese alloy had decreased by orders of magnitude, which makes the above-mentioned method feasible and applicable to the current process.

In addition to the dephosphorization of manganese ore, there are many findings on the separation of phosphorus and manganese from steelmaking slag, and many of the techniques are fundamentally common to those applied to manganese ore.

Shin et al. [8] proposed a process of selective separation of phosphorus and manganese in steelmaking slag by carbon reduction using the difference in chemical stability between  $\text{MnO}$  and  $\text{P}_2\text{O}_5$ . At the initial stage, steelmaking slag is carbon reduced under high-temperature and low-basicity ( $\text{CaO}/\text{SiO}_2$ ) conditions, and only phosphorus is selectively reduced and separated into a metal phase. If the remaining high-manganese–low-phosphorus slag is reduced together with iron ore, a high-manganese–low-phosphorus alloy can be produced. It should be noted that the conditions such as slag basicity and temperature are two main influencing factors for the selective reduction of phosphorus.

Kim [9] sulfurized the steelmaking slag to produce a liquid sulfide phase-matte, without phosphorus. A high-purity Fe–Mn alloy is produced by desulfurizing the matte. This method is based on the selective separation of the manganese sulfide phase. However, sulfur is an element that has a negative effect on steel quality, so a precise desulfurization process is required. Nevertheless, this manganese-rich sulfide phase is expected to be usable as a raw material for the production of low-phosphorus manganese.

Teratoko et al. [10], Numata et al. [11], and Du [12] focused on phosphorus's property to be distributed unevenly in a specific phase in steelmaking slag, which is easily soluble in a weakly acidic aqueous solution. A selective leaching method of the phosphorus-containing phase was proposed that consisted of immersing it in an aqueous solution. The authors reported that a high phosphorus separation rate was achieved.

Jung et al. [13,14] measured the manganese distribution between the  $\text{CaO-SiO}_2\text{-FeO-MnO}$  slag and liquid silver in equilibrium with solid Fe, in the range of a MnO concentration up to 20%. It was reported that the Mn distribution ratio between the slag and metal increased linearly with the FeO content. The authors also investigated the effect of  $\text{CaF}_2$ , which increased the MnO and FeO activities.

Morales and Fruehan [15] measured the activity of MnO and FeO in  $\text{CaO-SiO}_2\text{-FeO-MnO-P}_2\text{O}_5\text{-MgO}$  slags with MnO concentrations up to 25%. They reported that the activity of MnO showed a positive deviation from the ideal solution and increased with rising basicity, while the FeO activity decreased.

Ohta et al. [16] measured the MnO activity coefficient at 1873 K in a  $\text{MnO-SiO}_2\text{-Al}_2\text{O}_3$  slag at mullite and alumina saturation using a slag–metal equilibration technique. The authors derived an iso-activity diagram where the MnO activity did not substantially depend on the  $\text{SiO}_2/\text{Al}_2\text{O}_3$  ratio, but only on the MnO concentration.

Woo [17] measured the MnO activity in an  $\text{MnO-SiO}_2\text{-Al}_2\text{O}_3$  slag by equilibrating it with a Pt–Mn alloy and confirmed that the MnO activity was not influenced by the  $\text{SiO}_2/\text{Al}_2\text{O}_3$  ratio but increased significantly with the increasing MnO mole fraction above 0.5.

While there have been many measurements of MnO activity conducted in steel slags under conditions relevant to the steelmaking process, such as high temperature and low MnO content, fewer studies have focused on high-MnO slags containing more than 30% MnO and conditions applicable to the selective reduction process of Mn ore. According to

the work of Shin [8], in the process of the carbon reduction of manganese ore, phosphorus oxide is easily reduced in the strongly reducing atmosphere and can then be removed from the slag phase. However, the optimal conditions for suppressing manganese reduction require further investigation. Moreover, thermodynamic knowledge of manganese behavior in high manganese slag under carbothermic reduction process conditions is insufficient, and the influence of temperature requires more examination [18].

The present study investigates the optimal slag composition and conditions for manganese reduction control in the slag phase during selective reduction. MnO activity was measured in MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> ternary slags at different temperatures to predict the manganese distribution ratio based on the activity value of the MnO. The results derived serve to better understand the thermodynamic properties and behavior of Mn in high-manganese slags.

## 2. Materials and Methods

### 2.1. Principles of MnO Activity Measurement

The chemical equilibrium method was applied to measure the MnO activity in the ternary MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> slag by equilibrating it with molten silver. The equilibrium is governed by Equation (1):



Using the equilibrium constant  $K_1$  of the reaction (1), the MnO activity in the slag can be expressed as follows:

$$a_{\text{MnO}} = \gamma_{\text{MnAg}} \cdot X_{\text{MnAg}} \cdot P_{\text{O}_2}^{\frac{1}{2}} \cdot K_1 \quad (2)$$

where  $\gamma_{\text{MnAg}}$  is the activity coefficient of manganese in silver,  $X_{\text{MnAg}}$  is the molar fraction of manganese in silver, and  $P_{\text{O}_2}$  is the oxygen partial pressure. All activities are based on pure substances. The manganese activity coefficient in molten silver in Equation (2) and the equilibrium constant in Equation (1) have been derived in previous studies as functions of temperature  $T$  (see Equations (3) and (4)). Since the oxygen partial pressure was controlled, the activity of manganese in slag can be calculated by measuring the molar fraction of manganese in silver after the equilibrium experiment.

$$\log \gamma_{\text{MnAg}} = -\frac{358}{T} - 0.0178 \quad (3)$$

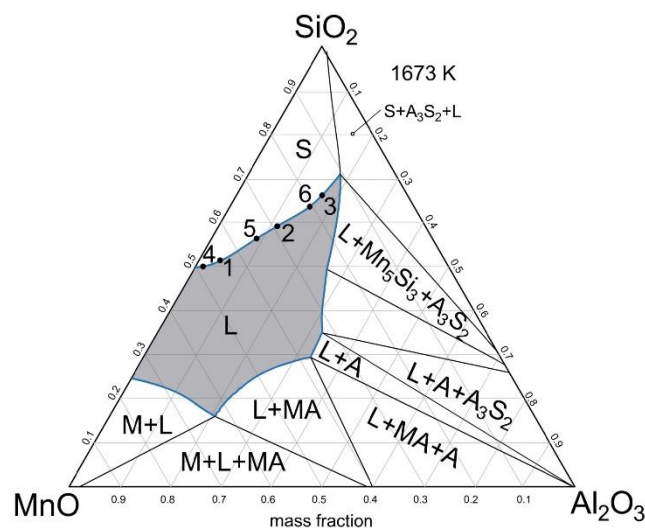
$$\log K_1 = \frac{21234}{T} - 4.6 \quad (4)$$

### 2.2. Apparatus and Experiment

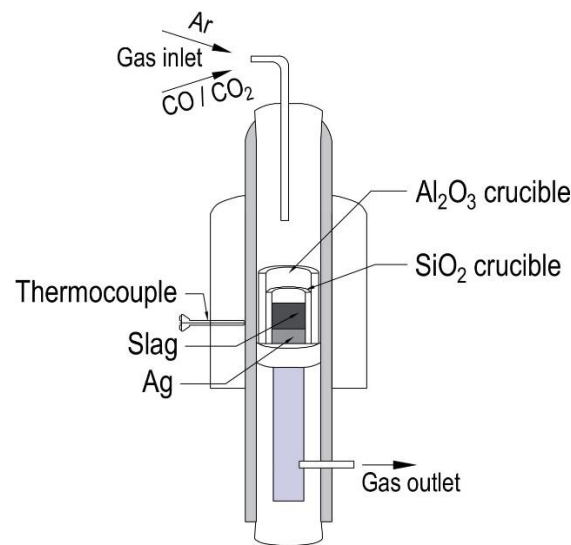
Slags with compositions from 1 to 6 were prepared by mixing reagent grade powders of MnO, SiO<sub>2</sub>, and Al<sub>2</sub>O<sub>3</sub> in a predetermined stoichiometric ratio (see Table 1). The prepared slag compositions correspond to the liquidus line plotted on the ternary diagram by FactSage™ 7.3 (Thermfact/CRCT (Montreal, Canada) and GTT-Technologies (Aachen, Germany)). (see Figure 1) [19]. SiO<sub>2</sub> crucibles were used for the equilibration experiment (Fujiwara Seisakusho, Tokyo, Japan, with outer diameter 11 mm, inner diameter 9 mm, and height 30 mm). It is assumed that the uptake from the crucible was minimal and the slag composition during the experiment did not change. The prepared slags were SiO<sub>2</sub>-saturated, and thus, SiO<sub>2</sub> presented in excess with respect to the liquid phase, which allowed significant erosion of the crucible by the slag to be prevented.

**Table 1.** Slag compositions used in the equilibration experiment (mass %).

Slag	MnO	SiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>
1	46	51	3
2	33	58	9
3	26	59	15
4	49	46	5
5	36	53	11
6	25	61	14

**Figure 1.** MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> phase diagram showing slag compositions and the liquidus line at 1673 K calculated by FactSage™ 7.3, (L—liquid slag, M—monoxide, MA—spinel, A—corundum, A<sub>3</sub>S<sub>2</sub>—mullite, S—cristobalite) [19].

The prepared slag samples were brought into equilibrium with metallic silver under controlled atmospheres using a CO/CO<sub>2</sub> gas mixture in the range of temperatures 1623–1723 K. The experimental temperatures were above the melting temperature of silver, which is 1240 K [20]. Experiments with slags from 1 to 3 were performed at three different temperatures, while slags 4, 5, and 6 were only equilibrated at 1673 K. The time to reach equilibrium was confirmed in a preliminary experiment and the holding time was set to 48 h. The silver used was a granular sample with 99.9% purity. The SiO<sub>2</sub> crucible containing 2 g of slag sample and 5 g of metallic silver was placed into an Al<sub>2</sub>O<sub>3</sub> protection crucible (Nikkato, Osaka, Japan, outer diameter 38 mm, height 45 mm) and introduced into the vertical electric resistance furnace (NEMS, BE1700T, Saitama, Japan). The sample was then heated up to the desired temperature under Ar gas atmosphere first, and after reaching the target temperature, the CO/CO<sub>2</sub> gas mixture was switched on, starting the equilibration at this point. The experimental apparatus is shown schematically in Figure 2. The furnace setup comprises a molybdenum disilicide heating element, a reaction tube (Nikkato, Osaka, Japan, outer diameter 50 mm, inner diameter 42 mm, length 1000 mm), a thermocouple that was connected to the furnace body, and a PID (proportional integral derivative) controller (Toei Kagaku Sangyo, SU12N5383-MANNNN, Sendai, Japan). In the equilibrium experiment, a precise adjustment of the temperature is critical, which is why temperature measurements were also performed before the equilibration. A thermocouple was inserted from the upper part of the reaction tube, and temperature measurements were carried out at each position. The temperature range in the hot zone of the furnace was determined to be within  $\pm 3$  K.



**Figure 2.** Schematic diagram of the experimental equipment.

The partial pressure of oxygen ( $P_{O_2}$ ) in the furnace was kept constant at  $10^{-12}$  by controlling the mixing ratio of the CO and CO<sub>2</sub> gases during the experiment. The following equilibrium reaction between gases is considered for the determination of the oxygen partial pressure [21]:



$$\Delta G^\circ = -565160 + 172.03 T \quad (6)$$

$$\log K_5 = \frac{14673}{T} - 4.46 \quad (7)$$

The equilibrium constant  $K_5$  of Equation (7) is expressed as follows:

$$K_5 = \frac{P_{CO_2}}{P_{CO} \cdot P_{O_2}^{\frac{1}{2}}} \quad (8)$$

and the oxygen partial pressure  $P_{O_2}$  is derived as follows:

$$P_{O_2} = \left(\frac{P_{CO}}{P_{CO_2}}\right)^2 \cdot \frac{1}{K_5} \quad (9)$$

During the experiment, the partial pressure ratio of the CO and CO<sub>2</sub> gases was determined by means of the flow ratio mixing method. The concentration of each component of the mixed gas was obtained by controlling the flow rate of each gas. Therefore, by setting the predetermined ratio of the flow rate, the partial pressures of CO and CO<sub>2</sub> gases and the oxygen partial pressure could be controlled. The gas flow was measured using a small mass flow controller (Kofloc, FCC-3000, Nagoya, Japan). A precision membrane flowmeter (HORIBA, VP-2U, Kyoto, Japan) was used to calibrate the flow measurement unit. After 48 h, the fused slag and silver samples were quickly withdrawn from the furnace and quenched with 99.9% pure argon gas stream. The slag and metal samples were separated. Then, the slag was embedded into the resin and polished with silicone carbide water-resistant abrasive paper, with the addition of 14,000 μm diamond paste. The obtained sample was supplied for the EPMA (electron probe microanalyzer) analysis performed by the (JEOL, JXA-8200, Tokyo, Japan) microanalyzer.

The silver sample was carefully cleaned from the fused slag residues and washed with ethanol. The clean Ag sample was then dissolved in strong HNO<sub>3</sub> acid. The manganese concentration, which then came from the silver sample into the resulting solution, was measured by ICP-AES (inductively coupled plasma atomic emission spectroscopy,) analysis using an ICP emission spectrometer (SPECTRO ARCOS EOP system, METEK, Kleve, Germany).

### 3. Results

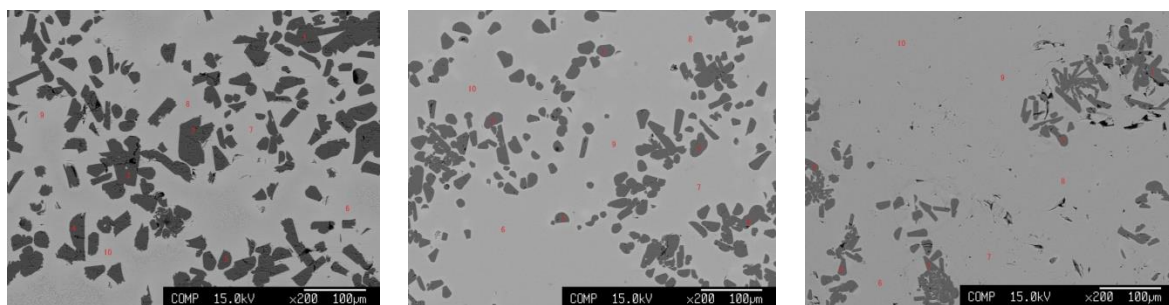
#### *Equilibrium between Ag and the MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> Slag*

The Mn content in the Ag phase after equilibrium at each experimental temperature measured by ICP-AES analysis is shown in Table 2.

**Table 2.** Mn concentrations in silver measured by ICP-AES after equilibrium in the temperature range 1623–1723 K.

Sample	Mn in Ag, ppm		
	1673 K	1723 K	1623 K
Slag 1	720.61	1193.85	1896.13
Slag 2	584.77	465.21	609.52
Slag 3	343.27	350.48	116.77
Slag 4	1150	-	-
Slag 5	850	-	-
Slag 6	350	-	-

The compositions of the original slags from 1 to 6 (see Table 1) derived after the equilibration experiment and EPMA analysis were renamed as slag 1-a, 2-a, 3-a, etc., respectively. The analyzed compositions are shown in Table 3 and represent the average measurement result obtained from multiple points from the slag matrix, avoiding the precipitated SiO<sub>2</sub>. The images of the investigated 1-a, 2-a, and 3-a slags are shown as examples in Figure 3. It can be seen that the SiO<sub>2</sub> phase is distributed throughout the slag matrix, as all slags were saturated with silicon dioxide.



**Figure 3.** EPMA images of slags 1-a (left), 2-a (middle) and 3-a (right).

Table 3 lists the measured chemical compositions of the silver phase and slags after equilibrium and the calculated values of the MnO activity and activity coefficients based on the experimental data, the regular solution model (RSM), and FactSage™ calculations. The concentration of components such as Mn in silver was considered to be sufficiently low.

The activity of manganese oxide  $a_{\text{MnO}}$  was calculated by employing Equation (2), using the determined oxygen partial pressure and the equilibrium composition of the silver and slag phase. The activity coefficient of MnO was obtained by dividing the calculated activity value by the MnO mole fraction  $X_{\text{MnO}}$  in the slag, which was derived by chemical analysis.

Table 3. Slag and metal compositions after equilibrium.

Slag No	T, K	CO <sub>2</sub> /CO Flow Ra- tio	Mn in Ag Mass %	Slag Composition, EPMA, Mass %			Experiment		RSM		FactSage™	
				MnO	SiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	a <sub>MnO</sub>	γ <sub>MnO</sub>	a <sub>MnO</sub>	γ <sub>MnO</sub>	a <sub>MnO</sub>	γ <sub>MnO</sub>
1-a	1673	0.021	0.072	48.62	48.00	3.38	0.105	0.25	0.161	0.356	0.130	0.28
2-a		0.021	0.058	34.10	56.30	9.60	0.085	0.26	0.071	0.222	0.080	0.25
3-a		0.021	0.034	25.03	59.33	15.64	0.050	0.20	0.040	0.157	0.040	0.17
1-b	1723	0.012	0.119	49.42	47.46	3.12	0.075	0.18	0.169	0.400	0.14	0.30
2-b		0.012	0.047	33.93	56.91	9.16	0.029	0.09	0.081	0.253	0.087	0.25
3-b		0.012	0.035	24.14	61.61	14.25	0.022	0.09	0.045	0.180	0.040	0.17
1-c	1623	0.039	0.191	49.70	47.30	3.00	0.666	1.57	0.133	0.314	0.13	0.31
2-c		0.039	0.061	34.70	55.80	9.50	0.214	0.67	0.062	0.193	0.08	0.24
3-c		0.039	0.012	25.50	60.00	14.50	0.041	0.17	0.034	0.145	0.02	0.08
4-a	1673	0.021	0.115	49.00	47.90	3.10	0.167	0.37	0.196	0.431	0.14	0.31
5-a		0.021	0.085	38.90	51.50	9.60	0.123	0.34	0.121	0.331	0.10	0.33
6-a		0.021	0.033	29.30	54.10	16.60	0.050	0.17	0.029	0.107	0.03	0.10

#### 4. Discussion

According to the experimental results, in a high-manganese slag with an MnO concentration of about 49%, the MnO activity value was 0.1 at 1673 K. A decrease in the MnO concentration led to a drop in the MnO activity values. This trend is shown in Figure 4, where the activity of the MnO is plotted against the MnO mole fraction in the slag at each temperature of the experiment. The decrease in the MnO concentration becomes more evident when the MnO mole fraction drops below 0.4. The trend agrees with the results of Woo et al., who investigated the MnO activity in a similar ternary MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> slag at 1823 K [17].

Figure 5 (left) depicts the effect of the (%MnO)/(%SiO<sub>2</sub>) ratio in the slag on the activity of the MnO in the temperature range of 1623–1723 K. The values of a(MnO) increase gradually with the increasing (%MnO)/(%SiO<sub>2</sub>) ratio. The effect of temperature is shown in Figure 5 (right), where the logarithms of the MnO activity values were plotted against the inverse temperature. As seen in the plot, the MnO activity decreased with temperature. Since the activity value the slag 1-c at 1623 K appeared to be very high, it was neglected in the plots. In addition, the investigated slag compositions were plotted on a ternary phase diagram of the MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> system, and the relationship between the MnO activity, activity coefficient, and slag composition at each point is shown (see Figure 6). The liquidus line on the diagram was calculated by FactSage™ 7.3 at 1673 K.

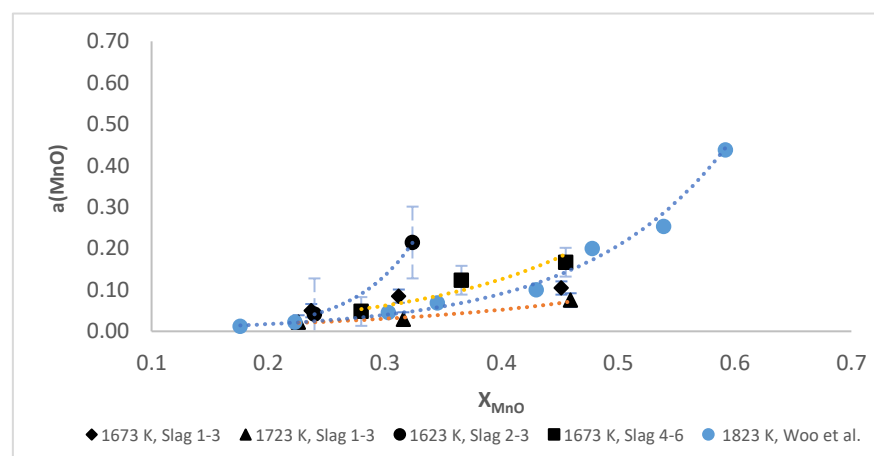
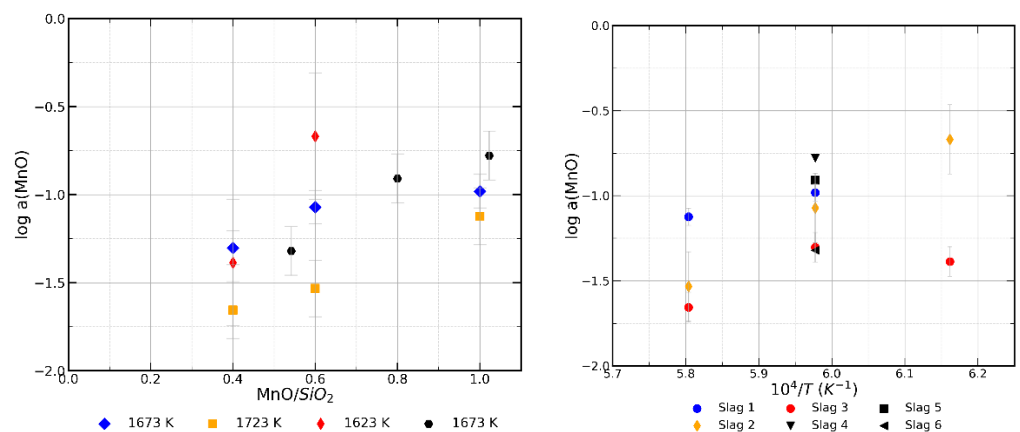
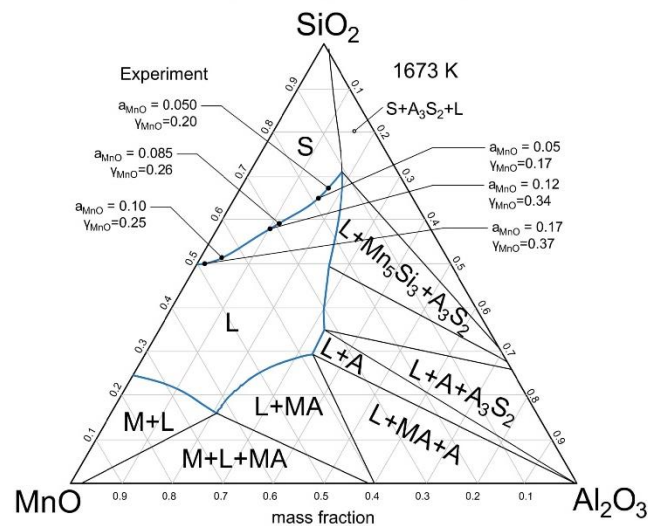


Figure 4. The dependence of the MnO activity on the MnO mole fraction. Comparison of the results obtained in this work with those of Woo et al. [17].



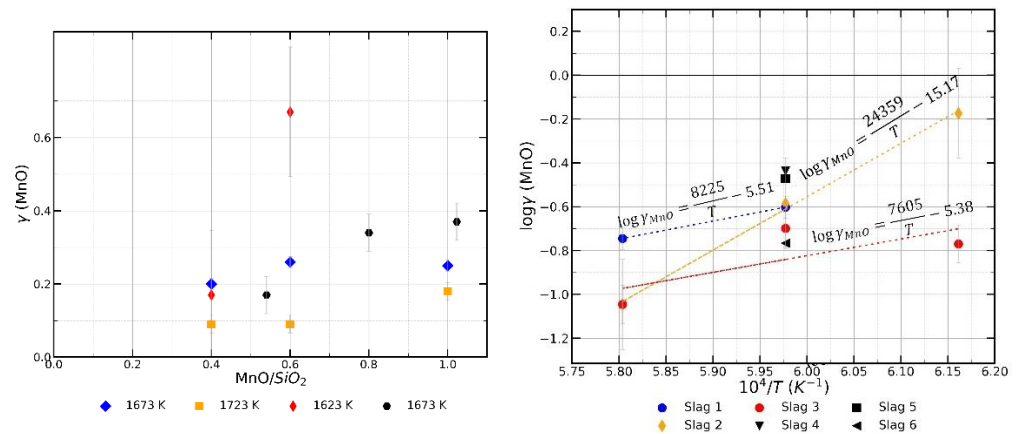
**Figure 5.** The relationship between the MnO activity of the liquid phase and the MnO/SiO<sub>2</sub> ratio at each temperature (**left**) and between the MnO activity and temperature in each slag composition (**right**).



**Figure 6.** MnO activity and activity coefficient of SiO<sub>2</sub>-saturated liquid phase at 1673 K, (L—liquid slag, M—monoxide, MA—spinel, A—corundum, A<sub>3</sub>S<sub>2</sub>—mullite, S—cristobalite).

It can be seen that the rate of the decrease in the activity coefficient is moderate in comparison to the decrease in the MnO concentration. Based on the experimental data, the relationship between the activity coefficients of the MnO and the slag composition is presented graphically in Figure 7 (left). The activity coefficient values rise with the increasing (%MnO)/(%SiO<sub>2</sub>) ratio; however, the rate of the increase is sharper at the lower temperature of 1623 K. As a next step, the derived temperature dependence for each (%MnO)/(%SiO<sub>2</sub>) ratio of 0.4, 0.6, and 1 was plotted in Figure 7 (right) and shows that the results vary with respect to the slag composition used.

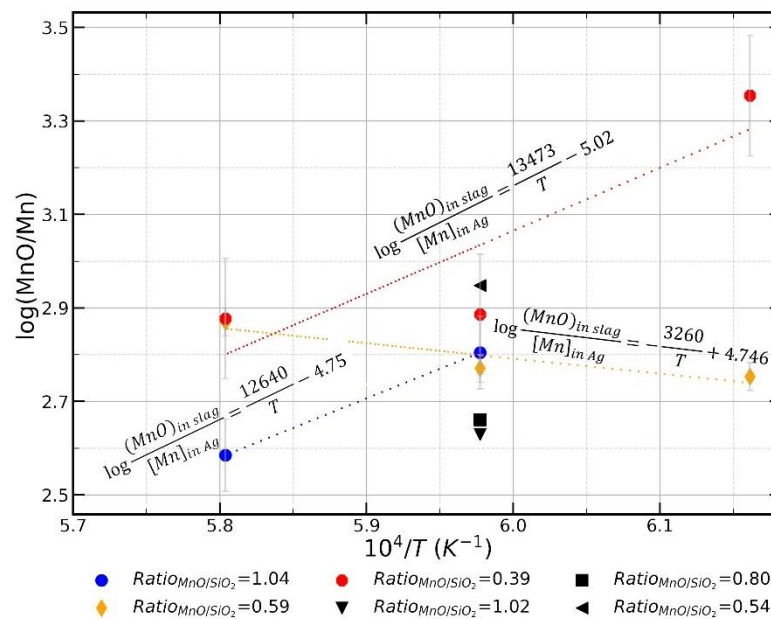




**Figure 7.** The relationship between the MnO activity coefficient and the MnO/SiO<sub>2</sub> ratio at each temperature (left) and between the MnO activity coefficient and temperature in each slag composition (right).

The reaction of Mn oxidation is exothermic; therefore, the equilibrium manganese distribution is expected to decrease with rising temperatures.

The manganese distribution ratio at each temperature derived from experimental data is plotted in Figure 8. A strong influence of the slag composition on the equilibrium distribution is evident.



**Figure 8.** The relationship between  $\log (\% \text{MnO})/[\% \text{Mn}]$  in Ag and inverse temperature.

#### Thermodynamic Assessment of $a_{\text{MnO}}$ and $\gamma_{\text{MnO}}$

The RSM was employed to calculate the activity and activity coefficient of MnO in the molten slag on the basis of interaction energies between cations and to compare with experimentally derived values.

A regular solution is a solution in which molecules are randomly arranged during mixing, but their interaction forces are not equal. Considering this, the activity coefficients of the slag components are calculated as a function of temperature and composition. The activity coefficient of component *i* in a multicomponent regular solution can be derived from the following quadratic formalism [22]:

$$RT\ln\gamma_i = \sum_j \alpha_{ij}X_j^2 + \sum_j \sum_k (\alpha_{ij} + \alpha_{ik} - \alpha_{jk})X_jX_k + I \quad (10)$$

where  $R$  is the gas constant;  $T$  is the temperature in Kelvin;  $\alpha_{ij}$  is the interaction energy between the cation of component  $i$  and component  $j$ ;  $X_j$  is the mole fraction of component  $j$  and  $I$  is the conversion factor. Since the reference state of component  $i$  stands for a hypothetical liquid oxide of a regular nature in Equation (10),  $\gamma_i$  for a real slag can be obtained using a conversion factor for the activity coefficient of a hypothetical liquid oxide, and that of a real solution in a multicomponent slag system.

The values of the interaction energies necessary for consideration in this calculation are given in Table 4. By applying Equation (10) for the currently investigated MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> slag system and using the conversion factor  $I$  at 1723 K, Equation (11) is obtained:

$$RT\ln\gamma_{MnO} = -75,310X_{SiO_2}^2 - 83,680X_{Al_2O_3}^2 - 31,380X_{SiO_2}X_{Al_2O_3} - 12,574 \text{ (J mol}^{-1}\text{)} \quad (11)$$

**Table 4.** Interaction energies between cations of slag components used in the experiment (J) [22].

Ion-ion	Mn <sup>2+</sup>	Si <sup>4+</sup>	Al <sup>3+</sup>
Mn <sup>2+</sup>	-	-75,310	-83,680
Si <sup>4+</sup>	-75,310	-	-127,610
Al <sup>3+</sup>	-83,680	-127,610	-

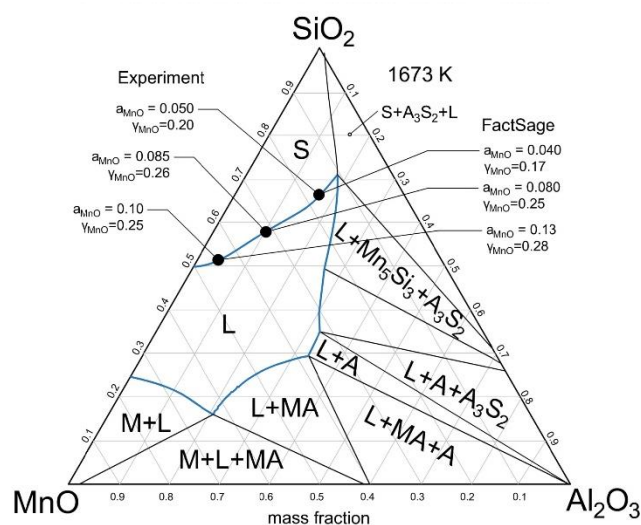
The results obtained using the RSM for the MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> slag are listed in Table 3. The lowest deviation between the calculated and measured values of the activity coefficients  $\gamma_{MnO}$  is at 1673 K and the activity values are generally in agreement at this temperature. However, at 1723 K, the significant deviation between both the activities and activity coefficients is evident for low and high-MnO slags. At the lowest temperature of 1623 K, the calculated and measured values of the activity coefficients are about 3 to 5 times different for high-MnO slags, for example, slags 1-c and 2-c, while the results agree on the low MnO-side for slag 3-c.

As the next step, the calculations were performed using the thermodynamic computation software FactSage™ 7.3. The contrast between the measured and calculated values follows the same tendency as in the case of the RSM application, e.g., at temperatures 1623 and 1732 K. Moreover, the highest deviations are at 1623 K for the same slags 1-c and 2-c, while at 1673 K, the results are for the most part in good agreement.

It can also be concluded that the calculations carried out with the RSM and FactSage™ 7.3 correspond with each other. The values of the MnO activity and activity coefficient calculated with FactSage™ 7.3 at 1673 K are presented graphically in Figure 9.

Careful observation of the results suggests that the most significant contrast between the experimentally measured and calculated values (comparing with both RSM and FactSage™) occurs at 1623 K, in particular where the highest activity value was derived (e.g., 0.21 for the slag 2-c). This agrees with Woo et al.'s statement that the application of the RSM and the conversion factor proposed by Ban-ya can only represent the system adequately for activity values below 0.15 [17,22]. Although there is scattering between the measured and calculated values of the activity coefficients, the RSM can still be used. However, not all compositional ranges can be sufficiently covered. As a result, in this case, modification of the conversion factor may be needed.

It is evident that the representation of the thermodynamic behavior of the entire real slag system covering different compositional ranges could not be conducted by applying thermodynamic models only. The combination of experimental investigation along with the thermodynamic models and FactSage™ is useful.



**Figure 9.** MnO activity and activity coefficient of the SiO<sub>2</sub>-saturated liquid phase at 1673 K, (L—liquid slag, M—monoxide, MA—spinel, A—corundum, A<sub>3</sub>S<sub>2</sub>—mullite, S—crystalite).

## 5. Conclusions

It is necessary to understand the thermodynamic properties of MnO in the slag for better process control during carbothermic reduction. Moreover, knowledge regarding manganese behavior is required for other metallurgical processes dealing with high-manganese slags. Thus, investigations of influencing factors such as the slag composition and temperature should be conducted. In this work, the activity and activity coefficient of MnO were experimentally measured in the ternary MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> slag saturated with SiO<sub>2</sub> on the basis of the equilibrium Mn content in the molten silver under a controlled gas atmosphere at 1623, 1673 and 1723 K. The following findings were obtained:

- The MnO activity values measured at 1673 K range from 0.05 to 0.10 and rise with the increasing MnO concentration from 25 to 48%.
- Both the MnO activity and activity coefficients rise with the increasing MnO/SiO<sub>2</sub> ratio at all temperatures in the experiments.
- The MnO activity coefficient exhibits a tendency to decrease with temperature in all slag compositions. Based on the derived results, it was shown that the temperature dependence of the MnO activity coefficient differs depending on the slag composition. For each MnO/SiO<sub>2</sub> ratio, that is, 0.4, 0.6, and 1, the temperature relation was derived and presented graphically.
- For the MnO equilibrium distribution ratio between silver and the slag phase, the derived temperature dependence was found to be influenced by the slag composition to a great extent.
- The results show that both the temperature and the slag composition are important parameters that influenced the results.
- The thermodynamic assessment of the MnO activities and activity coefficients was carried out by applying the RSM and FactSage™. Comparing the experimental results with the RSM, the values deviate from those calculated at 1623 K the most. The slags with the highest MnO activity, particularly higher than 0.15, showed the largest scattering in activity coefficients from the model calculations. At 1673 K, the contrast was the lowest.
- The deviation between the measured and calculated values may be due to high slag acidity with respect to the applied conversion factor, corresponding to more basic slags. In this case, the correction of a conversion factor may be needed.

- The MnO activity and activity coefficient values were plotted on the MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> ternary diagram at 1673 K and compared with those calculated using FactSage™ 7.3. At 1673 K the values generally agree, while at 1623 K, the deviation is the greatest. The scattering between the experimental values and FactSage™ calculations follows the same trend as in the case of the RSM application. The calculated by RSM and FactSage™ values are in good agreement.
- The findings derived in this work contribute to a better understanding of thermodynamic properties of manganese in the MnO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> slag and provide better process control in the metallurgical processes involving slags containing high MnO.

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